

Form Approved
OMB No. 0704-0188

Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing this collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden to Department of Defense, Washington Headquarters Services, Directorate for Information Operations and Reports (0704-0188), 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302. Respondents should be aware that notwithstanding any other provision of law, no person shall be subject to any penalty for failing to comply with a collection of information if it does not display a currently valid OMB control number. **PLEASE DO NOT RETURN YOUR FORM TO THE ABOVE ADDRESS.**

Standard Form 298 (Rev. 8-98)
Prescribed by ANSI Std. Z39.18

MEMORANDUM FOR PRS (In-House Publication)

FROM: PROI (STINFO)

23 May 2001

SUBJECT: Authorization for Release of Technical Information, Control Number: **AFRL-PR-ED-TP-2001-122**
Boatz, Jerry, et al., "New Materials Design"

DoD High Performance Computing Users Conference
(Biloxi, MS, 18-22 June 2001) (Deadline: 15 June 2001)

(Statement A)

1. This request has been reviewed by the Foreign Disclosure Office for: a.) appropriateness of distribution statement, b.) military/national critical technology, c.) export controls or distribution restrictions, d.) appropriateness for release to a foreign nation, and e.) technical sensitivity and/or economic sensitivity.

Comments: _____

Signature _____ Date _____

2. This request has been reviewed by the Public Affairs Office for: a.) appropriateness for public release and/or b) possible higher headquarters review.

Comments: _____

Signature _____ Date _____

3. This request has been reviewed by the STINFO for: a.) changes if approved as amended, b.) appropriateness of references, if applicable; and c.) format and completion of meeting clearance form if required

Comments: _____

Signature _____ Date _____

4. This request has been reviewed by PR for: a.) technical accuracy, b.) appropriateness for audience, c.) appropriateness of distribution statement, d.) technical sensitivity and economic sensitivity, e.) military/national critical technology, and f.) data rights and patentability

Comments: _____

APPROVED/APPROVED AS AMENDED/DISAPPROVED

PHILIP A. KESSEL Date
Technical Advisor
Space and Missile Propulsion Division

New Materials Design

Jerry Boatz¹, Mark S. Gordon², Gregory Voth³, Sharon Hammes-Schiffer⁴, and
Ruth Pachter⁶

¹Air Force Research Laboratory, AFRL/PRSP, Edwards AFB, CA 93524

²Department of Chemistry, Iowa State University, Ames, IA 50011

³Department of Chemistry and Henry Eyring Center for Theoretical Chemistry, University of
Utah, Salt Lake City, UT 84112

⁴Department of Chemistry, The Pennsylvania State University, University Park, PA 16802

⁵Air Force Research Laboratory, AFRL/MLPJ, Wright-Patterson AFB, OH 45433

Abstract. In this paper we report our recent results on the design of materials with controlled properties by the application of computational chemistry methods, new algorithms and scalable software.

1 Introduction

Developments of new materials (such as new high energy density materials; POSS (polyhedral oligomeric silsesquioxanes) that are highly resistant to extreme environments and are therefore desirable coatings for rocket engines; nonlinear optical materials; and liquid crystals) are important in many applications of interest to the Air Force. However, the design of large and complex material systems with controlled properties is challenging. Therefore, the calculations we describe in this paper would not have been possible without scalable quantum chemistry codes such as GAMESS (General Atomic and Molecular Electronic Structure System), developed under the auspices of a DoD CHSSI (Computational High Performance Scalable Software Initiative) grant, Gaussian, and molecular dynamics codes, as well as the availability of the DoD Computation Centers through our computational challenge project.

2 Results and Discussion

2.1 High-Energy Density Materials (HEDM)

2.1.1 Bond dissociation energies: a comparative study of theoretical methods^{1,2}

The purpose of this study is to perform a systematic comparison of several theoretical methods for computing the bond dissociation energies (BDEs) of a prototypical energetic compound. The

¹ This work was done in collaboration with Donald L. Thompson and Dan C. Sorescu, Department of Chemistry, Oklahoma State University, Stillwater, OK 74078.

² We thank Dr. Jeffrey A. Sheehy for several stimulating discussions.

computational methods include single-reference (Moller-Plesset second order perturbation theory (MP2)[1], Gaussian-2 using MP2 (G2(MP2))[2], and coupled cluster with single and double substitutions (CCSD(T))[3]), multiconfigurational (complete active space self-consistent field (CASSCF)[4] and multiconfiguration quasi-degenerate perturbation theory (MCQDPT)[5]), as well as density functional theory (DFT, using Becke's three parameter hybrid exchange functional with the Lee, Yang, and Parr correlation (B3LYP)[6]) methods. The energetic compounds of interest are 1,1-diamino-2,2-dinitroethylene (FOX-7) and its simpler analogues. FOX-7 was recently synthesized[7] and found to have superior impact and friction sensitivity properties relative to the well-known HMX and RDX explosives. This molecular system is an example of a "push-pull" ethylene compound, which has both electron donating (NH_2) and electron withdrawing (NO_2) groups attached to a common ethylene linkage. Such compounds are potentially stabilized through " π " resonance delocalization and hence may possess non-negligible multiconfigurational character. Therefore, this molecule should be a good testbed for comparing DFT and single-configuration methods against multiconfigurational techniques.

The geometries of the reactants and products of the C- NH_2 and C- NO_2 homolytic bond dissociation reactions of FOX-7 and its analogues have been fully optimized and verified as local minima at the B3LYP, MP2 and CASSCF levels, using the 6-311G(d,p) basis set.[8] CCSD(T)/MP2, MCQDPT/CASSCF, and G2(MP2) energies were also computed. The putative resonance delocalization of the " π " electrons in FOX-7 suggests that in addition to the (2e,2o) C-N σ, σ^* space needed to describe homolytic bond cleavage, an appropriate active space for the MCSCF calculations should include the (2e,2o) C- $\pi + \pi^*$, (4e,3o) NO_2 $\pi + \text{non-bonding} + \pi^*$, and (2e,1o) NH_2 lone pair electrons and orbitals, described in further detail elsewhere.³ The degree of multiconfigurational character is assessed by calculation of the natural orbital occupation numbers (NOONs) of the MP2, CCSD(T), and CASSCF first order density matrices. As shown in a recent work by Gordon[9] et al., the occurrence of "non-physical" MP2 and CCSD(T) NOONs is an indication that a single configuration treatment may be inadequate. All calculations were performed using GAMESS[10], GAUSSIAN98[11], and ACESII.[12]

At the MP2 level, aminoethylene is the only closed shell compound without at least one non-physical NOON. Although some calculations are still in progress, it appears that all open-shell species except the amino and nitro radicals have multiple non-physical NOONs at this level of theory. In contrast, none of the compounds have non-physical NOONs at the CCSD(T) level, although several open shell CCSD(T) calculations are still in progress. This indicates that CCSD(T) is able to capture more of the non-dynamical correlation than MP2 and hence should be more reliable. The CASSCF NOONs show that the lone pair orbital on the NH_2 group has little multiconfigurational character. However, the nominally unoccupied NO_2 π^* (0.08–0.14 electrons) orbitals have relatively large populations, suggesting that the C- NO_2 BDEs may not be as well described by single reference methods as the C- NH_2 BDEs.

The predicted C- NH_2 BDEs for aminoethylene (which include zero-point energy corrections) are all within 9 kcal/mol of the experimental value of 102 kcal/mol[13] and range from 93 (MCQDPT) to 108 (MP2) kcal/mol. The B3P86/6-31+G(d,p)[14] and G2(MP2) predictions are both within about 1 kcal/mol of experiment. MP2 overestimates the BDE by 6 kcal/mol and MCQDPT underestimates it by about 9 kcal/mol, with CCSD(T) (97.4 kcal/mol) and B3LYP

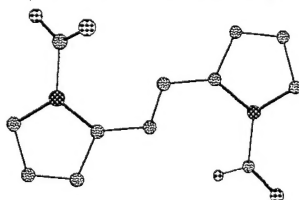
³ Boatz, J.A., Sorescu, D.C., Thompson, D.L.: manuscript in preparation

(96.7 kcal/mol) yielding intermediate values. CCSD(T) is the single reference method in closest agreement with MCQDPT.

For 1-amino-2-nitroethylene, MP2 again gives the largest C-NH₂ BDE (126 kcal/mol) and MCQDPT the smallest (112 kcal/mol). CCSD(T) is again the single-reference method in closest agreement with MCQDPT. Although the C-NH₂ BDE calculations for the entire set of ethylene derivatives are incomplete, it appears that (a) MP2 consistently gives the highest C-NH₂ BDEs and (b) the MCQDPT BDEs are in good agreement with CCSD(T). In contrast, the DFT BDEs show no clear patterns of behavior. For nitroethylene, the C-NO₂ BDEs range from 59.5 kcal/mol (MCQDPT) to 79.4 kcal/mol (G2(MP2)). CCSD(T) is the single-reference method, which is in closest agreement with MCQDPT, although the difference between these two values is greater than in the case of the C-NH₂ BDEs of aminoethylene. The B3LYP value of 68.8 kcal/mol is in good agreement with CCSD(T) (70.9 kcal/mol), although the B3P86 prediction of 74.5 kcal/mol[14] is closest to MP2 (77.7 kcal/mol). Similar observations apply to the predicted C-NO₂ bond energies in 1-amino-2-nitroethylene, which range from 69.3 (MCQDPT) to 88.8 kcal/mol (G2(MP2)). CCSD(T) (78.2 kcal/mol) is the single-reference method in closest agreement with the MCQDPT result and also is in good agreement with the B3LYP value (81.4 kcal/mol). MCQDPT also predicts the smallest C-NO₂ BDE for 1-amino-2-nitroethylene (69 kcal/mol). Although the C-NO₂ BDE calculations are not yet complete, it appears that (a) G2(MP2) and MP2 consistently give the highest C-NO₂ BDEs and (b) the single reference method in best agreement with MCQDPT is CCSD(T).

2.1.2 High-energy density materials

A recent major focus centers around a series of proposed monopropellants⁴, such as:

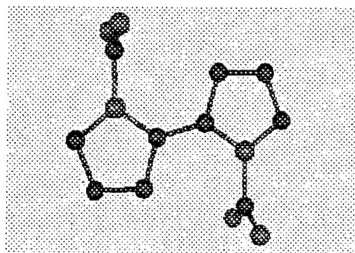


I_{sp} , which is a measure of the energy released by the reaction of the fuel divided by the mass of the products, is of primary interest:

$$I_{sp} \propto [\Delta H/m]^{1/2} \quad (1)$$

Using a combination of isodesmic reactions and the G2 model, the heat of formation for this species is predicted to be 456.8 kcal/mol. This translates to an I_{sp} of 240 sec, compared with 230 sec for hydrazine, thus being a very promising fuel. However, it is important to consider the stability of such high-energy species to various reactions, before asserting their viability as fuels. One possible reaction is the loss of molecular nitrogen to form the smaller species shown below:

⁴ Work of the late Rob Schmitt and Jeff Botaro, SRI.



Using our highly scalable second order perturbation theory code with a large basis set, we have found that this decomposition process is exothermic by 60 kcal/mol. It may be that the barrier for decomposition is large, in which case the species may still be viable. The search for the transition state for this reaction is in progress.

Recently the first new all-nitrogen compound in nearly a century, N_5^+ [15], was synthesized. The counterion for this species was AsF_6^- , the result being a rather unstable solid. In order to understand the potential energy surface for this species, we performed MP2 calculations on the various stationary points. The first step in this process was to separate the two ions by a large distance and optimize the geometry. The result was an ion pair $[AsF_6^-][N_5^+]$ complex that is, not surprisingly, 98 kcal/mol below the separated ions. What was somewhat surprising is that this complex is 47 kcal/mol below the neutral species that is formed when one F^- ion is transferred to N_5^+ to form neutral $AsF_5 + FN_5$. The former is a well-known species, but the latter has not previously been reported. FN_5 has several isomers, whose relative energies and the barriers separating them were determined using large basis set CCSD(T) calculations. The lowest lying isomers were found to have small energy barriers (< 10 kcal/mol) separating them from the much more stable decomposition products $FN_3 + N_2$. The exothermicity of this decomposition is more than 50 kcal/mol, so the products $AsF_5 + FN_3 + N_2$ are slightly lower in energy than the original ion pair $[AsF_6^-][N_5^+]$. Since FN_3 is itself an energetic species with a relatively small barrier to further decomposition, it is therefore not surprising that the original species made by Christe et al. is unstable.

2.1.3 Quantum molecular dynamics (MD) simulations of low temperature HEDM

Traditional MD algorithms can be extended to incorporate equilibrium quantum mechanical effects through the use of discretized Feynman path integrals[16]. The resulting algorithm, path integral MD (PIMD), is highly amenable to parallelization because each physical particle becomes rigorously mapped onto a collection of classical-like quasiparticles at different values of imaginary time, each connected to its neighbor by a harmonic-like force. The discretized PIMD algorithm can be load-balanced across computational nodes because the quasiparticles with the same imaginary time index experience identical forces. The inter-node communications are minimal since there are only nearest-neighbor harmonic forces to compute. The PIMD algorithm, no matter how complex the physical system, can always be made to achieve excellent scalability and thus high performance. Moreover, classical MD techniques such as constant temperature and pressure algorithms can be readily incorporated into PIMD. The algorithm is even amenable to a two-tier level of parallelism, the first being over the imaginary time slices (quasiparticles), the second over the force loop for the interparticle interactions in large (thousand or even million) particle simulations.

In an important theoretical advance[17,18,19,20], PIMD methods have been extended to include quantum dynamical effects (i.e., not just equilibrium properties). This computational

approach is called "Centroid Molecular Dynamics" (CMD). It incorporates the dominant quantum dynamical effects of a many-body system into a classical-like MD framework, thus significantly extending the range and applicability of MD methods. The basic result of CMD[19-22] is that the dynamical correlations of quantum particles in a general many-body system can be accurately computed by running classical-like trajectories on an effective potential which includes the effects of quantum zero point energy and tunneling. The task of integrating the CMD equations is not trivial since the effective potential is a quantum potential of mean force, requiring an "on the fly" dynamical averaging at each timestep. A number of powerful algorithms for solving the CMD equation have been developed[22]. The best CMD algorithm so far is the "hyperparallel" CMD approach[21] using an adiabatic algorithm[22] in the first tier of parallelism to determine the effective potential; a second tier of parallelism can be introduced in which, for each of the averaging slaves, the MD steps are sent to additional nodes for each discretized path integral imaginary time slice[23]. A third tier of parallelism for very large systems can be utilized for each imaginary time slice, in which the classical force loop is split over several nodes. The scaling of the hyperparallel CMD code on the IBM SP and SGI O2000, as well as a Pentium Linux cluster, is excellent (Fig. 1). The parallelization of the long range forces (Ewald summation) was also achieved.

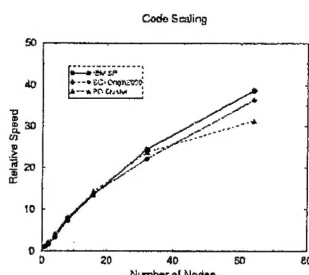


Fig. 1. Scaling of CMD

The PIMD and CMD algorithms are of general interest in all areas of condensed matter computer simulation. However, they are particularly important for the HEDM program in the computational study of, e.g., solid hydrogen, helium, or nitrogen matrices which contain isolated reactive species such as lithium, boron, aluminum, hydrogen atoms, organic radicals, etc. Such matrices are high priority targets as possible HEDMs for Air Force space propulsion purposes. At the temperatures appropriate to the condensed phases of hydrogen or helium, the host matrix molecules will exhibit large quantum mechanical effects, but the treatment of these effects by a direct attack on the time-dependent Schrödinger equation is impossible. The key issue is the stability of the HEDM (i.e., the ability to trap the energetic impurities for some period of time and thus impede their diffusion and eventual recombination). Both the structural and dynamical aspects of the composite condensed phase HEDM systems are of interest. For the equilibrium structural studies, PIMD methods are adequate. For the dynamical studies, the CMD approach must be employed to directly simulate the diffusive and recombination steps of the guest atoms in the solid host. The long-term goal of this research is to develop general computational methods to rapidly and efficiently characterize proposed high energy density materials.

The structure and stability of atomic impurities trapped in solid *para*-hydrogen have been studied by employing large scale PIMD and CMD simulations at 4 K and zero external pressure. Starting from pure solid hydrogen, doped systems were prepared by substituting impurity atoms for hydrogen molecules at substitutional defect sites. The structural and thermodynamic

properties of the resulting HEDM were then calculated. In the case of atomic boron impurities, the impurity atom interacts anisotropically with the matrix *para*-hydrogen molecules because of its singly filled 2p orbital. To assess the effect of the electronic anisotropy in the potential, a comparison has been made between simulations in which the orientation-dependent (anisotropic) and orientation-averaged (spherical) B-H₂ potentials were used. In collaboration with the Alexander group[22], three matrices were investigated: (a) one with a single B atom site-substituted for a *p*-H₂ molecule, (b) one with a similar site-substituted matrix having a nearest-neighbor vacancy, (c) and one with a periodic "super cell" simulation box having two hydrogen bulk slab regions in order to study a B impurity near the *p*-H₂ surface.

It was found that significant "second-order" distortions of the lattice occur to permit an energetically favorable orientation of the 2p orbital, even in the absence of a neighboring vacancy. Especially when the B impurity is located near the surface, the spherically averaged potential provides a noticeably different description from the case of the anisotropic potential. These quantum MD simulations reveal the large structural differences that can be found in the lattice resulting from the anisotropic boron-hydrogen interaction potential. The implications of this behavior will be explored in future research.

2.2 POSS Compounds

There is great interest in POSS compounds because of their resistance to extreme environments and the ease with which they are synthesized. Since almost nothing is known experimentally about the mechanism by which they are formed, we have embarked on a long term project to determine the possible competing mechanisms as a function of catalyst, solvent, and substituents (e.g., in [23]). The synthesis begins with hydrolysis of RSiX₃ to RSi(OH)₃, followed by condensation to the siloxane RSi(OH)₂-O-Si(OH)₂R. Subsequent condensations lead to the 3D cage compounds. We have shown that the initial hydrolysis and condensation steps all have very high-energy barriers in the gas phase, but are being reduced to nearly zero by the presence of one water molecule added to represent the solvent. Thus, all steps leading to the initial disiloxanes and to the ring compounds D₃ and D₄ occur with net energy requirements of less than 10 kcal/mol. In this paper we describe current efforts to (a) determine substituent effects on these barriers, (b) determine the effects of adding additional water molecules, and (c) to compare the properties of the 3D POSS compounds with their Ti analogs. The substituents studied include R=H, methyl, t-butyl and phenyl, and X=Cl, OCH₃. There is experimental interest in synthesizing incompletely condensed POSS. Because the reaction is very fast, it always goes to completion. The interest in substituent effects lies in attempting to slow down the reaction, possibly using bulky R groups. There is industrial interest in the possibility that Ti POSS may be another class of materials with desirable properties. Hammes-Schiffer et al.[24] have been developing methodologies for the simulation of hydrogen transfer reactions, and most recently a new molecular orbital method for the simultaneous calculation of electronic and hydrogen vibrational wavefunctions, now incorporated into GAMESS. This method may be used to obtain minimum energy reaction paths and direct dynamics trajectories of hydrogen transfer reactions, with the advantage that nuclear quantum effects such as zero point energy and hydrogen tunneling are incorporated during the generation of the reaction paths and trajectories, rather than subsequently included as corrections. This approach will be combined with a mixed quantum/classical surface hopping method to study the quantum dynamics of hydrogen transfer reactions with an *ab initio* potential energy surface obtained "on the fly," and used to investigate

hydrogen transfer reactions in the hydrolysis and condensation steps required for the synthesis of POSS, in order to determine the influence of different trihalosilane reactants on the rates and yields of these steps and aid in the efficient synthesis of POSS.

2.3 Nonlinear Absorbing Materials

We have previously shown that density functional theory and the time-dependent density functional response theory (TDDFT) (B3LYP) accurately predict the structures and spectra of porphyrin and the larger tetraphenylporphyrin and its β -octahalogenated derivatives,[25,26] thus enabling their assessment as reverse saturable absorbing materials. For porphyrin (PH₂), the photoinduced tautomerization process has also been investigated[27] in order to discern between previously proposed discordant mechanisms. Recently we carried out calculations[28] for tetrabenzoporphyrin (ZnTBP) and phthalocyanine (ZnPc) (Fig. 2).

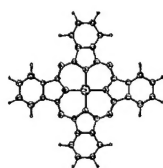


Fig. 2. Zinc phthalocyanine

Our calculations support the multiple bands in the B and higher energy regions of the ZnPc and ZnTBP spectra that have been determined from experiment. Despite considerable experimental and theoretical work on ZnPc and ZnTBP, their ground state electronic structures and UV spectra are not well understood. Previous theoretical studies provided only a qualitative understanding in the low-energy region, and none for the higher energy bands. We have elucidated significant shifts in Gouterman's four orbitals of ZnP upon tetraazasubstitution and tetrabenzoannulation.⁵ The benzo groups destabilize the $1a_{1u}$ orbital in ZnP, but they do not show a significant deviation from the four-orbital model for the interpretation of the Q and B bands of ZnTBP. However, tetraazasubstitutions considerably lower the $4a_{2u}$ orbital (HOMO-1) in ZnTBP. Also, it is shown that the near degeneracy of the HOMO and HOMO-1 in ZnP that provides the basis for Gouterman's four orbital model does not hold for ZnPc. In ZnPc, the near degeneracy of the HOMO-1 ($4a_{2u}$) with other occupied orbitals gives rise to a complex structure of the higher energy regions of the spectrum that were not reproduced by simple MO methods. The $4a_{2u}$ MO has significant π contributions from the benzo rings in ZnPc but not in ZnTBP. The Q and B bands are predicted to appear at 2.18 and 3.28 eV for ZnTBP, respectively, in excellent agreement with the maxima located at 2.02 and 3.18 eV as established from the absorption spectrum in a supersonic jet expansion, with similar values (2.02 & 3.10 eV) obtained from the absorption spectrum in an Ar matrix. Tetraazasubstitutions are predicted to redshift the Q band of ZnTBP by about 0.1 eV. The computed excitation energies in near quantitative agreement with experiment clearly provide a strong basis for the interpretation of the electronic spectra of such large molecular systems. TDDFT calculations are also being carried out for candidate two-photon absorption (TPA) materials,⁶ to be compared to those of Albota et al.[29]

⁵ Nguyen, K.A. and Pachter, R. J. Chem. Phys., manuscript in preparation

⁶ Day, P.N. and Pachter, R., work in progress

2.4 Liquid Crystals

Recent work is in progress for studying a liquid crystalline droplet to model the behavior of a liquid crystal in the bulk, with the details reported elsewhere.[30] Indeed, atomic level simulations are proven important in understanding the structure-property relations of materials. Due to the size of the LC droplet, very large-scale classical MD simulations of such systems are required, that may, become computationally prohibitive. The fast multipole method (FMM)[31], which uses a multiscale hierarchy of partitions of the volume and a divide-and-conquer strategy to compute the power series, allows all the forces to be computed to any specified accuracy in $O(N)$ operations. FMM3D (an implementation and improvement to FMM in 3-D), which contains a variety of schemes for computing multipole translations, has been implemented into the MD program that is being developed in our group.[32]

3 References

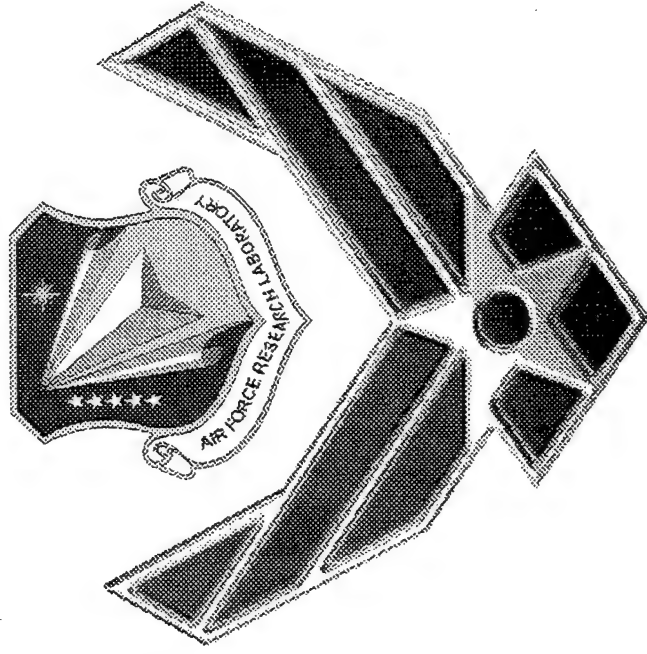
- 1 Pople, J.A., Binkley, J.S., Seeger, R., *Int. J. Quantum Chem.*, 10 (1976) 1; Bartlett, R.J., Silver, D.M.: *Int. J. Quantum Chem.*, 9 (1975) 183; Dupuis, M., Chin, S., Marquez, A.: In: *Relativistic and Electron Correlation Effects in Molecules*, Malli, G.: Ed. Plenum Press, NY 1994; Frisch, M.J., Head-Gordon, M., Pople, J.A.: *Chem. Phys. Lett.* 1660 (1990) 275-280
- 2 Curtiss, L.A., Raghavachari, K., Pople, J.A.: *J. Chem. Phys.* 98 (1993) 1293
- 3 See, for example, Bartlett, R.J., Stanton, J.F.: In: *Applications of Post-Hartree-Fock Methods: A Tutorial*, Reviews of Computational Chemistry, Vol. V, Lipkowitz, K.B. and Boyd, D.B., eds., VCH Publishers, Inc., New York, 1994
- 4 a) Siegbahn, P.E., Heiberg, A., Roos, B.O., Levy, B.: *Phys. Scr.* 21 (1980) 232. b) Roos, B.O., Taylor, P.R., Siegbahn, P.E.: *Chem. Phys.* 48 (1980) 157. c) Roos, B.O.: *Int. J. Quantum Chem.* S14 (1980) 175
- 5 Nakano, H.: *J. Chem. Phys.* 99 (1993) 7983-7992
- 6 Becke, A.D.: *J. Chem. Phys.* 98 (1993) 5648
- 7 Bemm, U., Östmark, H.: *Acta Cryst.* C54 (1998) 1997; Östmark, H., Langlet, A., Bergman, H., Wellmar, U., Bemm, U.: *Eleventh Int. Detonation Symp.*, 30 Aug – 4 Sep 1998, Snowmass Village, Colorado, USA
- 8 Krishnan, R., Binkley, J.S., Seeger, R., Pople, J.A.: *J. Chem. Phys.* 72 (1980) 650 (Six-component Cartesian d functions used.)
- 9 Gordon, M.S., Schmidt, M.W., Chaban, G.M., Glaesemann, K.R., Stephens, W.J., Gonzales, C.: *J. Chem. Phys.* 110 (1999) 4199
- 10 Schmidt, M.W., Baldridge, K.K., Boatz, J.A., Elbert, S.T., Gordon, M.S., Jensen, J.J., Koseki, S., Matsunaga, N., Nguyen, K.A., Su, S., Windus, T.L., Dupuis, M., Montgomery, J.A.: *J. Comput. Chem.* 14 (1993) 1347.
- 11 Gaussian 98, Revision A.7, Frisch, M. J., Trucks, G.W., Schlegel, H.B., Scuseria, Robb, M.A., Cheeseman, J.R., Zakrzewski, V.G., Montgomery, J.A., Jr., Stratmann, R.E., Burant, J.C., Dapprich, S., Millam, J.M., Daniels, A.D., Kudin, K.N., Strain, M.C., Farkas, O., Tomasi, J., Barone, V., Cossi, M., Cammi, R., Mennucci, B., Pomelli, C., Adamo, C., Clifford, S., Ochterski, J., Petersson, G.A., Ayala, P.Y., Cui, Q., Morokuma, K., Malick, D.K., Rabuck, A.D., Raghavachari, K., Foresman, J.B., Cioslowski, J., Ortiz, J.V., Baboul, A.G., Stefanov,

- B.B., Liu, G., Liashenko, A., Piskorz, P., Komaromi, I., Gomperts, R., Martin, R.L., Fox, D.J., Keith, T., Al-Laham, M.A., Peng, C.Y., Nanayakkara, A., Gonzalez, C., Challacombe, M., Gill, P.M.W., Johnson, B., Chen, W., Wong, M.W., Andres, J.L., Gonzalez, C., Head-Gordon, M., Replogle, E.S., and Pople, J.A.: Gaussian, Inc., Pittsburgh, PA, 1998
- 12 Stanton, J.F., Gauss, J., Watts, J.D., Lauderdale, W.J., and Bartlett, R.D.: *Int. J. Quantum Chem.*, S26 (1992) 879
- 13 Lias, S.G., Bartmess, J.E., Liebman, J.F., Holmes, J.L., Levin, R.D., and Mallard, W.G.: *J. Phys. Chem. Ref. Data* 17 (Suppl. 1) (1988) 1-872
- 14 Politzer, P., Concha, M.C., Grice, M.E., Murray, J.S., Lane, P., Habibollahzadeh, D.: *Theochem*, 452 (1998) 75
- 15 Christe, K.O., Wilson, W.W., Sheehy, J.A., and Boatz, J.A. *Angewandte Chemie*, 38 (1999) 2004
- 16 See, e.g., Berne, B.J., and Thirumalai, D.: *Annu. Rev. Phys. Chem.* 37 (1987) 401
- 17 Cao, J. and Voth, G.A.: *J. Chem. Phys.* 99 (1993) 10070
- 18 Cao, J. and Voth, G.A.: *J. Chem. Phys.* 100 (1994) 5106
- 19 Cao, J. and Voth, G.A.: *J. Chem. Phys.* 101 (1994) 6157
- 20 Cao, J. and Voth, G.A.: *J. Chem. Phys.* 101 (1994) 6168
- 21 Calhoun, A., Pavese, M., and Voth, G.A.: *Chem. Phys. Lett.* 262 (1996) 415
- 22 Krumrine, J.R., Jang, S., Alexander, M.H., and Voth, G.A.: *J. Chem. Phys.* 113 (2000) 9079
- 23 Kudo, T., Takako, and Gordon, M.S.: *J. Phys. Chem. A*, 104 (2000) 4058
- 24 Hammes-Schiffer, S.: *J. Phys. Chem. A* 102 (1998) 10443; Webb, S.P., Agarwal, P.K., and Hammes-Schiffer, S.: *J. Phys. Chem. B* 104 (2000) 888; Webb S.P., and Hammes-Schiffer, S.: *J. Chem. Phys.* 113 (2000) 5214
- 25 Nguyen, K. A., Day, P. N., and Pachter, R.: *J. Chem. Phys.*, 110 (1999) 9135; Nguyen, K. A., Day, P. N., and Pachter, R.: *J. Phys. Chem. A*, 103 (1999) 7378
- 26 Nguyen, K. A., Day, P. N., and Pachter, R.: *J. Phys. Chem.*, 104 (2000) 4755
- 27 Nguyen, K. A. and Pachter, R., *J. Phys. Chem.*, 104 (2000) 4549
- 28 Nguyen, K.A., and R. Pachter, *J. Chem. Phys.*, manuscript in preparation
- 29 Albota, M., Beljonne, D., Bredas, J.L., Ehrlich, J.E., Fu, J.Y., Heikal, A.A., Hess, S.E., Kogej, T., Levin, M.D., Marder, S.R., McCord-Maughon, D., Perry, J.W., Rockel, H., Rumi, M., Subramaniam, G., Webb, W.W., Wu, X.L., and Xu, C.: *Science* 281 (1998) 1653
- 30 Wang, Z., Patnaik, S., and Pachter, R.: *Computational and Theoretical Polymer Science*, to appear, 2001
- 31 Greengard, L., and Rokhlin, V.: *J. Comp. Phys.*, 73 (1987) 325
- 32 Wang, Z., Lupo, J.A., McKenney, A., and Pachter, R.: *Proceedings of SC99* (1999), November 13-19, 1999, Portland, OR

New Materials Design

DoD UGC, 18-22 Jun 01

Biloxi, MS

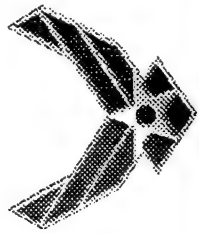


Jerry Boatz

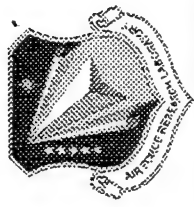
Research Chemist

Propulsion Directorate

Air Force Research Laboratory



NEW MATERIALS DESIGN



THE TEAM....

Prof. Mark S. Gordon

IOWA STATE UNIVERSITY

Prof. Gregory Voth

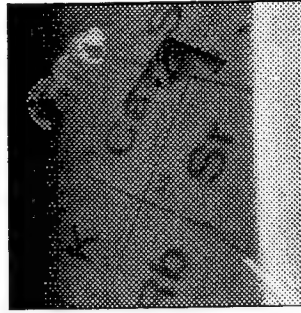


Prof. Sharon Hammes-Schiffer

PENNSTATE



Dr. Ruth Pachter, AFRL/MLPJ

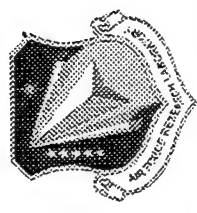


Dr. Jerry Boatz, AFRL/PRSP





OUTLINE



1. Project Overview

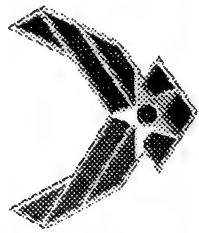
- High energy density matter
- Polyhedral oligomeric silsesquioxanes (POSS)
- Non-linear optical materials

2. Theoretical Methods and benchmarks

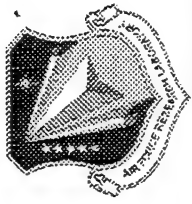
- Ab initio electronic structure theory
- Nuclear-electronic orbital approach
- Centroid Molecular Dynamics
- Fast Multipole MD (Ruth??)

3. Results

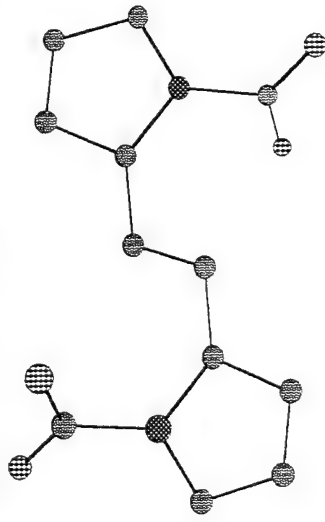
4. Summary



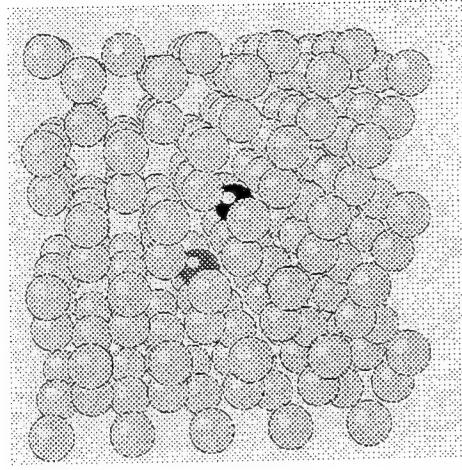
PROJECT OVERVIEW - HEDM



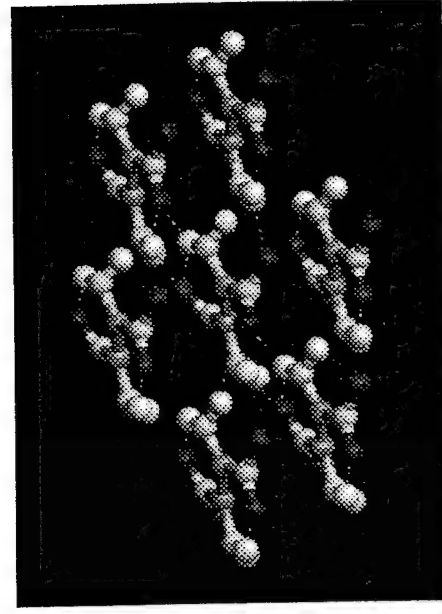
High Energy Density Matter -- next generation rocket propellants



High-nitrogen/polynitrogen compounds



Atom-doped solid hydrogen

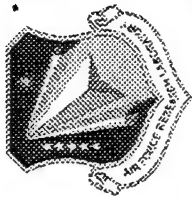


"FOX-7" (1,1-diamino-2,2-dinitroethylene)





PROJECT OVERVIEW - HEDM



Technical issues being addressed using CCM

1. High-nitrogen/polynitrogen compounds

Objective: identify, characterize, and synthesize stable compounds with high heats of formation, high densities

- structures, energy content, stabilities, reaction pathways

2. Energetic atoms in solid hydrogen

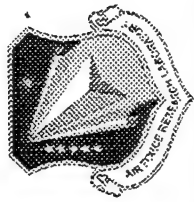
Objective: stabilize ~5% energetic atoms in solid hydrogen

- stabilities, mobilities, concentration limits of atoms stored in hydrogen matrices

3. "FOX-7"

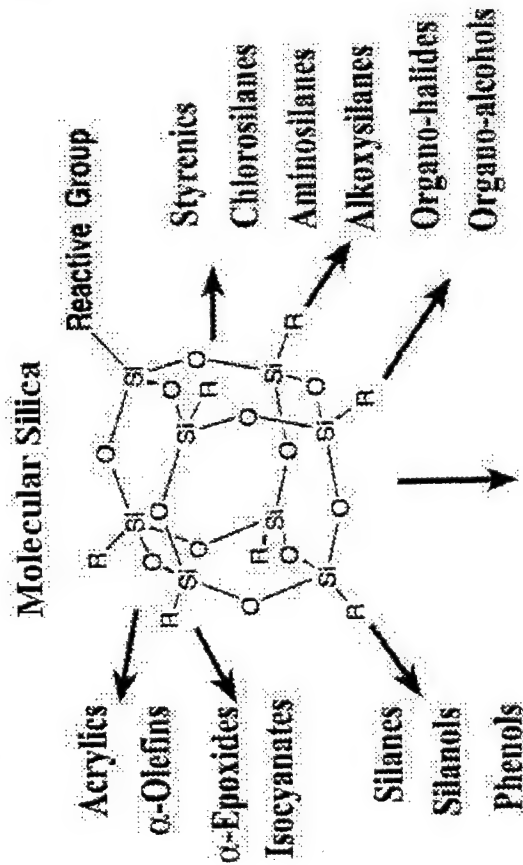
Objective: identify, characterize, and synthesize stable compounds with high heats of formation, high densities

- role of hydrogen bonding in reduced sensitivity, importance of non-additive inter/intramolecular interactions, decomposition mechanisms



PROJECT OVERVIEW - POSS

Polyhedral oligomeric silsesquioxanes -- next generation plastics



As Additives

Heat/abrasion resistant paints and coatings
Mechanical property/viscosity/thermal modifiers
Crosslinking agents
Fire retardants

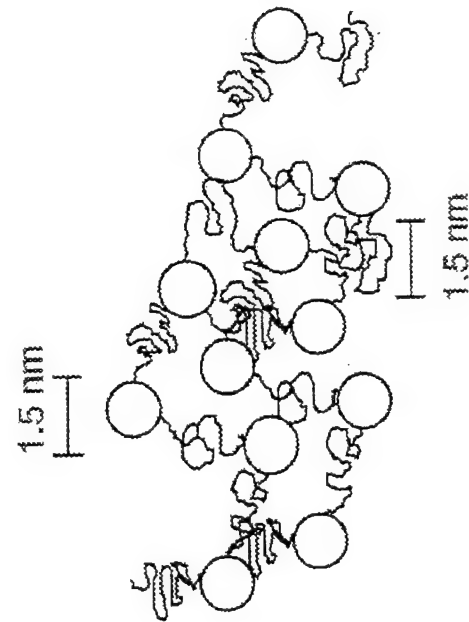
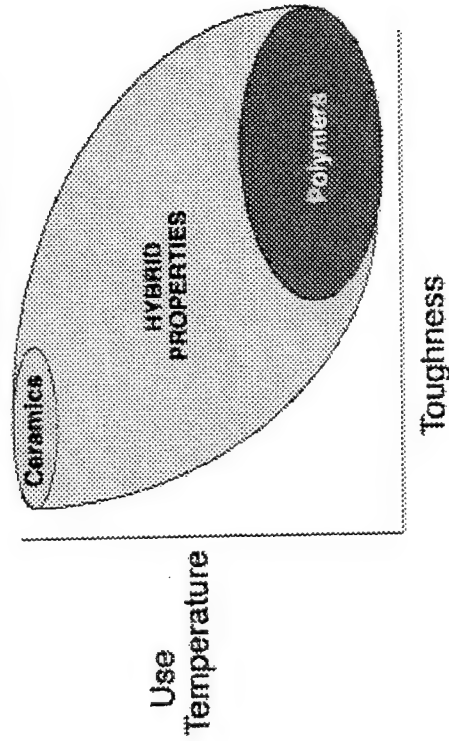
As Plastics

Medical materials
Space resistant resins
Packaging/coatings
Electronic materials
Optical Plastics

As Preceramics

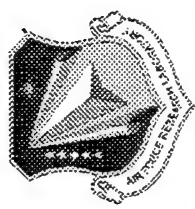
Ablative materials (nozzles, insulations etc.)
Claddings/electronics coatings
Precursors to glassy or ceramic matrices

HYBRID POLYMERS





PROJECT OVERVIEW - POSS



Technical issues being addressed using CCM

1. Mechanisms of formation

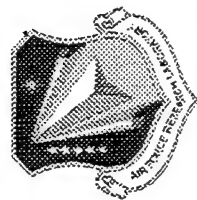
Objective: rational design and synthesis of POSS

- role of solvents, acid/base catalysis, substituent effects on mechanism of formation

2. Potential applications as molecular “sieves”

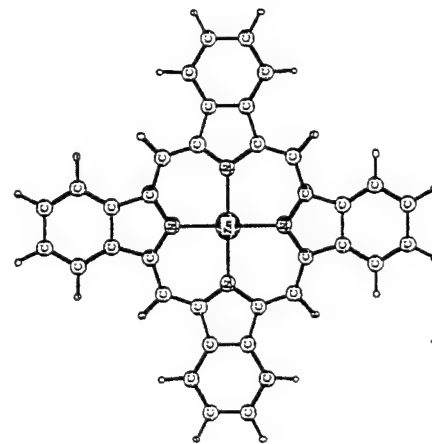
Objective: determine if POSS cages can be used to separate small molecules

- determine barriers to encapsulation of N_2 and O_2

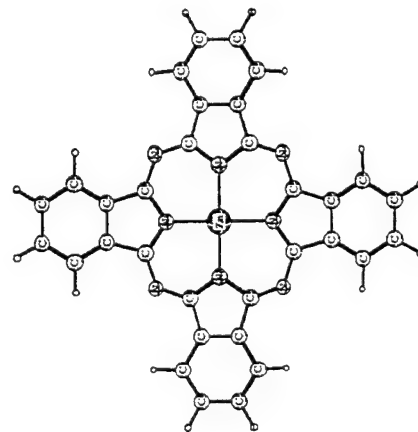


PROJECT OVERVIEW - NLO

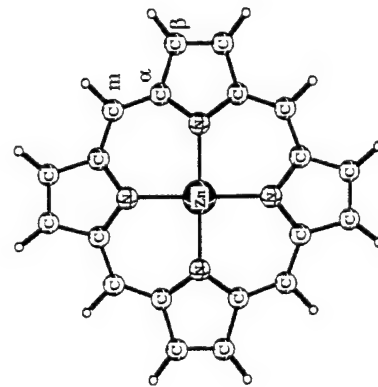
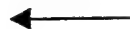
Non-linear optical materials for laser-hardened applications



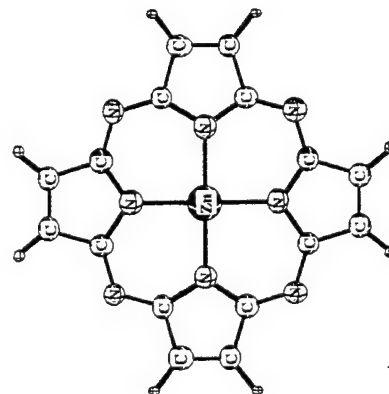
Azasubstitution



Benzannulation



Azasubstitution

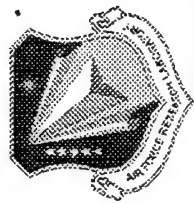


Benzannulation



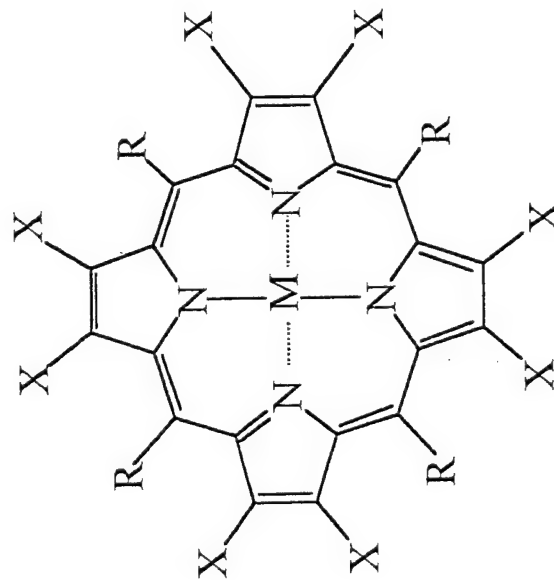


PROJECT OVERVIEW - NLO



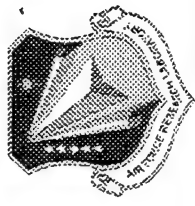
Technical issues being addressed using CCM

1. Mechanism of reverse saturable absorption (RSA)
2. "Tuning" of absorption spectrum by benzannulation, halide substitution



System	Property						
	M	X	R	IP	S0-S _n	S0-T ₁	T ₁ -T _n
PH ₂	H ₂	H	H	E, C	E, C	E, C	E, C
ZnP	Zn	H	H	E, C	E, C	E, C	E, C
TPPH ₂	Zn	H	ϕ	E, C	E, C	E, C	E, C
ZnTPP	Zn	H	ϕ	E, C	E, C	E, C	E, C
ZnTPPBr ₈	Zn	Br	ϕ	C	E, C	E, C	E, C

IP = Ionization Potential, S_0-S_n = Ground State Spectrum,
 S_0-T_1 = Singlet-Triplet Gap T_1-T_n = Triplet-Triplet Spectrum
 E = Experiment, C = Calculated



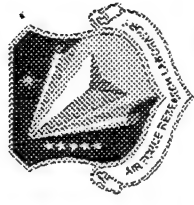
THEORETICAL METHODS

1. Ab initio electronic structure theory

- General Atomic and Molecular Electronic Structure System (GAMESS) -- a CHSSI code
- Nuclear-electronic orbital approach (NEO) for including nuclear quantum effects (important, e.g., in proton transfer reactions)

Various computational techniques are employed to solve the molecular electronic Schrödinger equation from quantum mechanics:

$$\left[-\frac{1}{2} \sum_i \nabla_i^2 - \sum_i \sum_{\alpha} \frac{Z_{\alpha}}{r_{i\alpha}} + \sum_i \sum_{j>i} \frac{1}{r_{ij}} \right] \Psi_{el} = E_{el} \Psi_{el}$$



THEORETICAL METHODS

1. Ab initio electronic structure theory (cont.)

- Most electronic structure codes use Born-Oppenheimer (i.e., “clamped nuclei”) approximation -- NOE method treats specified nuclei at QM level.

Nuclear-Electronic Hamiltonian

$$\begin{aligned}
 H_{\text{tot}}(\mathbf{r}_e, \mathbf{r}_q; \mathbf{r}_c) = & - \sum_i^{N_e} \frac{1}{2} \nabla_i^2 - \sum_i^{N_e} \sum_A^{N_c} \frac{Z_A}{r_{iA}} + \sum_i^{N_e} \sum_{j>i}^{N_e} \frac{1}{r_{ij}} \\
 & - \sum_I^{N_p} \frac{1}{2M_I} \nabla_I^2 + \sum_I^{N_p} \sum_A^{N_c} \frac{Z_A Z_I}{r_{IA}} + \sum_I^{N_p} \sum_{J>I}^{N_p} \frac{Z_I Z_J}{r_{IJ}} \\
 & - \sum_i^{N_e} \sum_I^{N_p} \frac{Z_I}{r_{iI}} + \sum_A^{N_c} \sum_{B>A}^{N_c} \frac{Z_A Z_B}{r_{AB}}
 \end{aligned}$$

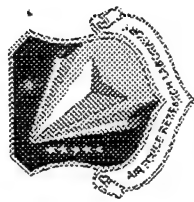
N_e : number of electrons (coordinates \mathbf{r}_e)

N_p : number of quantum nuclei (coordinates \mathbf{r}_p)

N_c : number of classical nuclei (coordinates \mathbf{r}_c)



THEORETICAL METHODS



Ab initio electronic structure theory

Current Status of parallel GAMESS

	<u>RHF</u>	<u>ROHF</u>	<u>UHF</u>	<u>GVB</u>	<u>MCSCF</u>
Energy	cdp	cdp	cdp	cdp	cdp
Analytic Gradient	cdp	cdp	cdp	cdp	cdp
Numeric Hessian	cdp	cdp	cdp	cdp	cdp
Analytic Hessian	cdp	cdp	-	cdp	-
MP2 energy	cdp	cdp	cdp	-	c
MP2 gradient	cdp	-	-	-	-
CI energy	cdp	cdp	-	cdp	cdp
CI gradient	cd	-	-	-	-

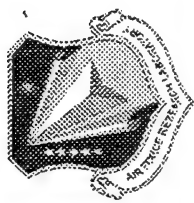
c = conventional disk storage of AO integrals

d = direct evaluation of AO integrals

p = run in parallel



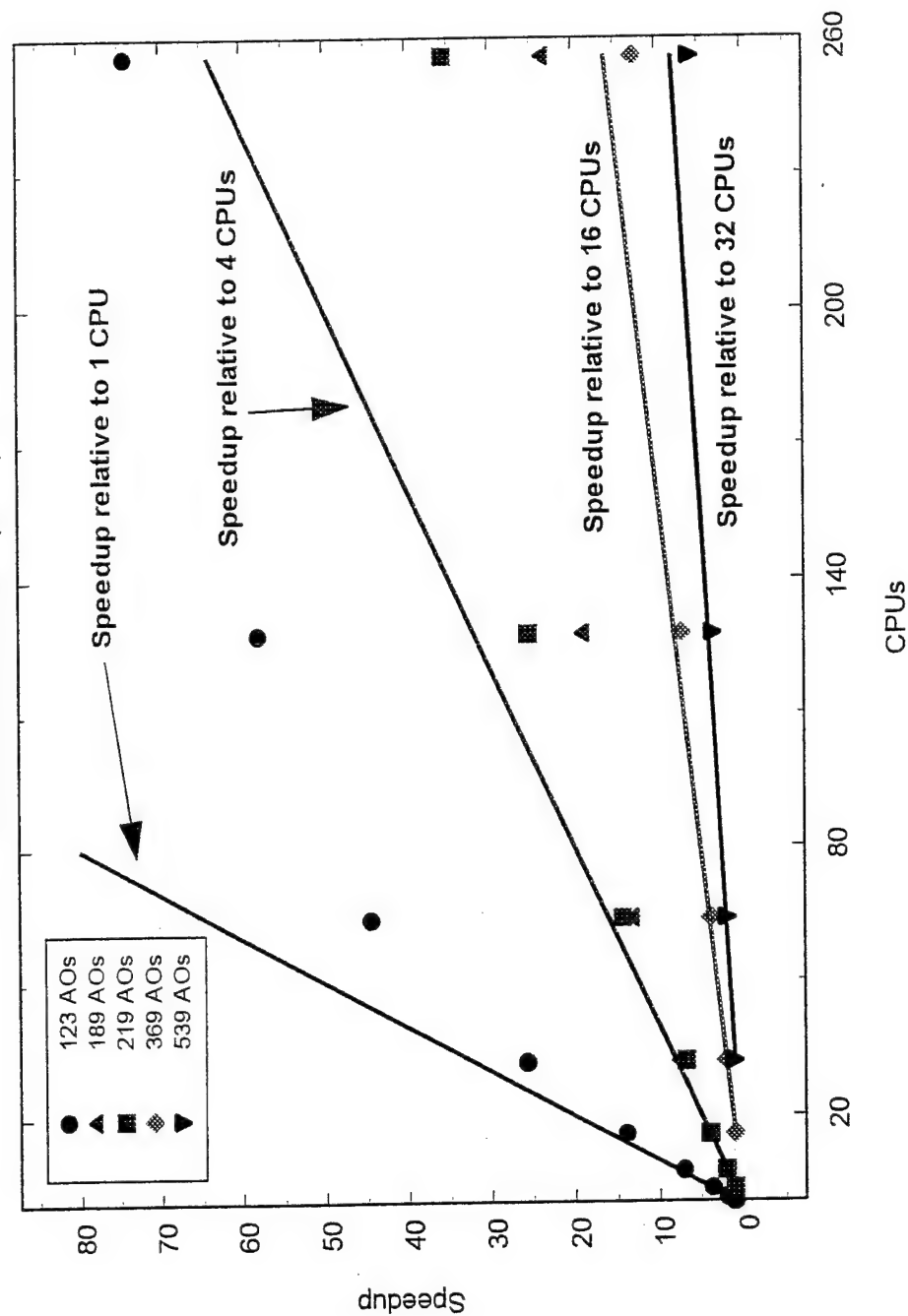
THEORETICAL METHODS

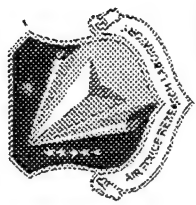


Ab initio electronic structure theory

MP2 Gradient Scalability Test

Silicocene molecule, $\text{Si}(\text{C}_5\text{H}_5)_2$

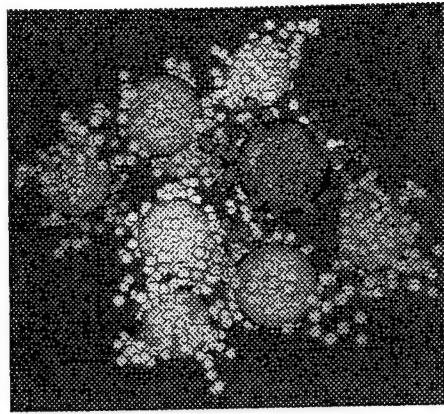
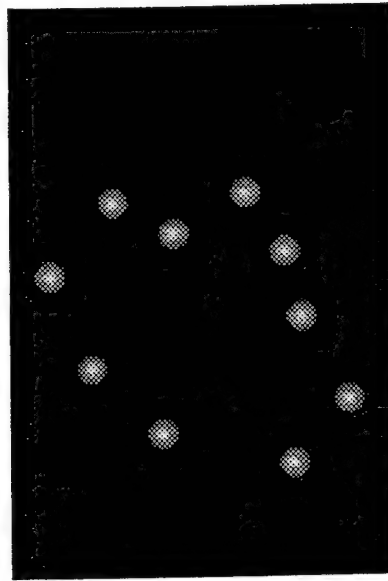
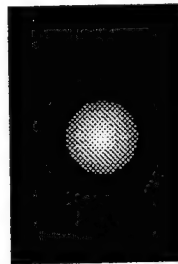




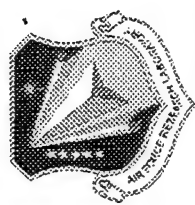
THEORETICAL METHODS

Path Integral Molecular Dynamics & Centroid Molecular Dynamics (CHSSI codes)

Simulation methods based on path integral techniques for mapping quantum particles onto “polymer ring” of classical quasiparticles:



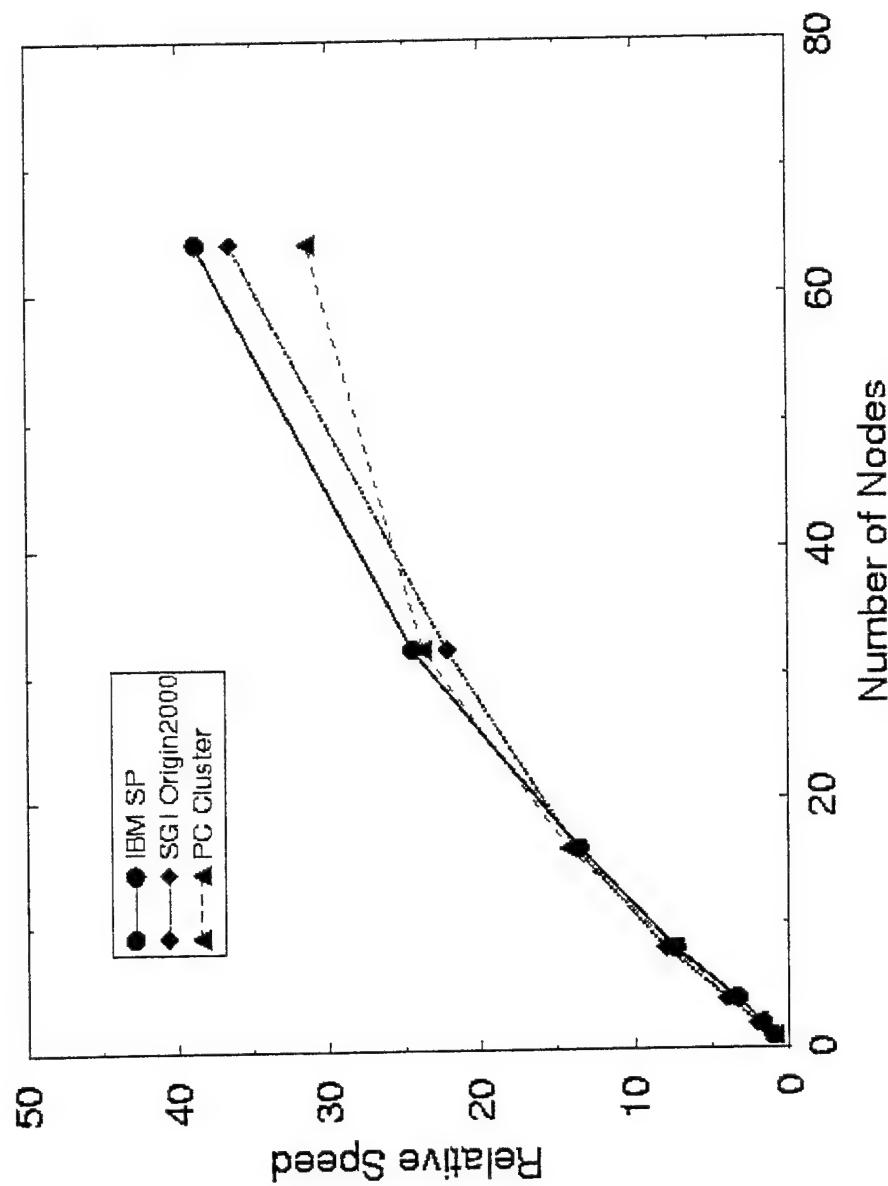
Each “real” particle is replaced by N ($50 < N < 500$) quasiparticles; classical dynamics done on collection of quasiparticles \Rightarrow natural, efficient parallelism.

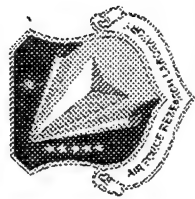


THEORETICAL METHODS

Path Integral Molecular Dynamics & Centroid Molecular Dynamics

Code Scaling

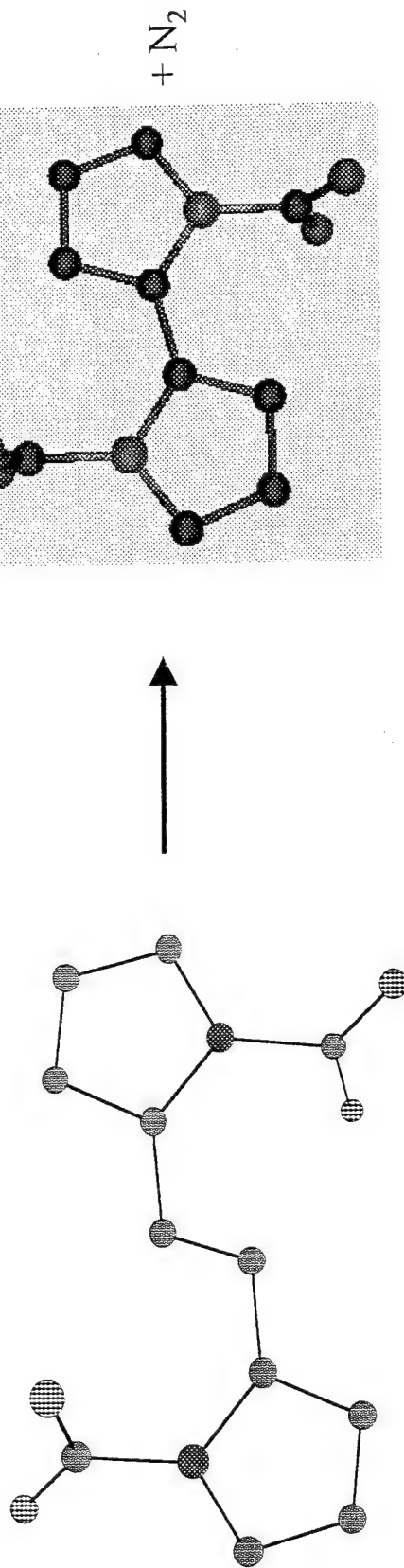




RESULTS - HEDM

High-nitrogen/polynitrogen compounds

Predicted $\Delta H_f = 457$ kcal/mol, $I_{sp} = 240$ sec
(I_{sp} for hydrazine = 233 sec)

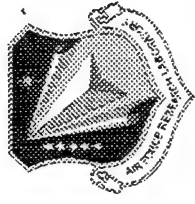


Predicted $\Delta H_r = -60$ kcal/mol, calculation of reaction barrier underway

Computational requirements: ~60,000 node-hrs, ERDC T3E, 64 GB memory



RESULTS - HEDM

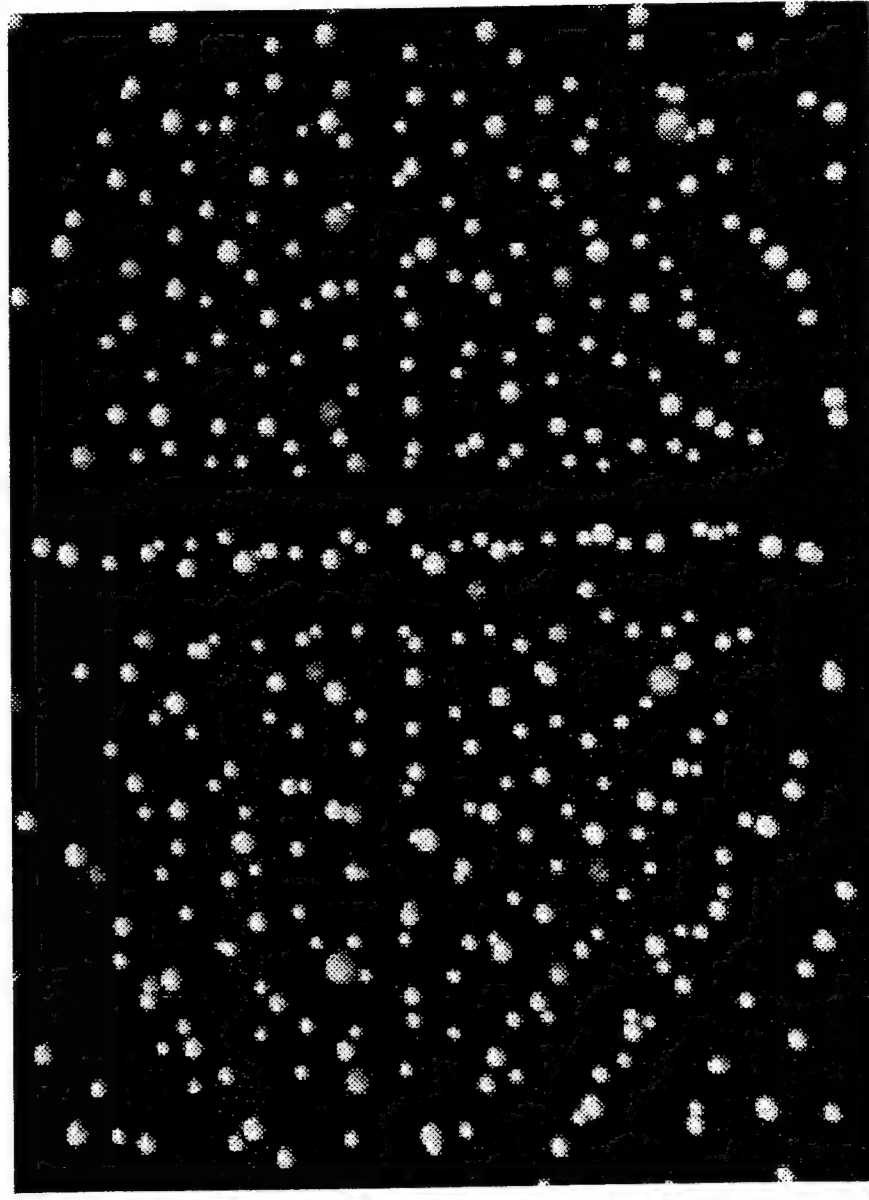


Atom-doped solid hydrogen

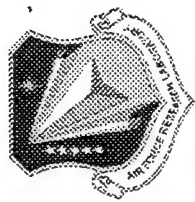
6.25% B atoms in solid *para*-H₂

Previous key results

1. In sH₂, B atoms more stable than Li atoms
2. No recombination of B atoms seen at concentrations up to 6.25%.
3. "Forced" recombination of B atoms does not trigger phase separation.



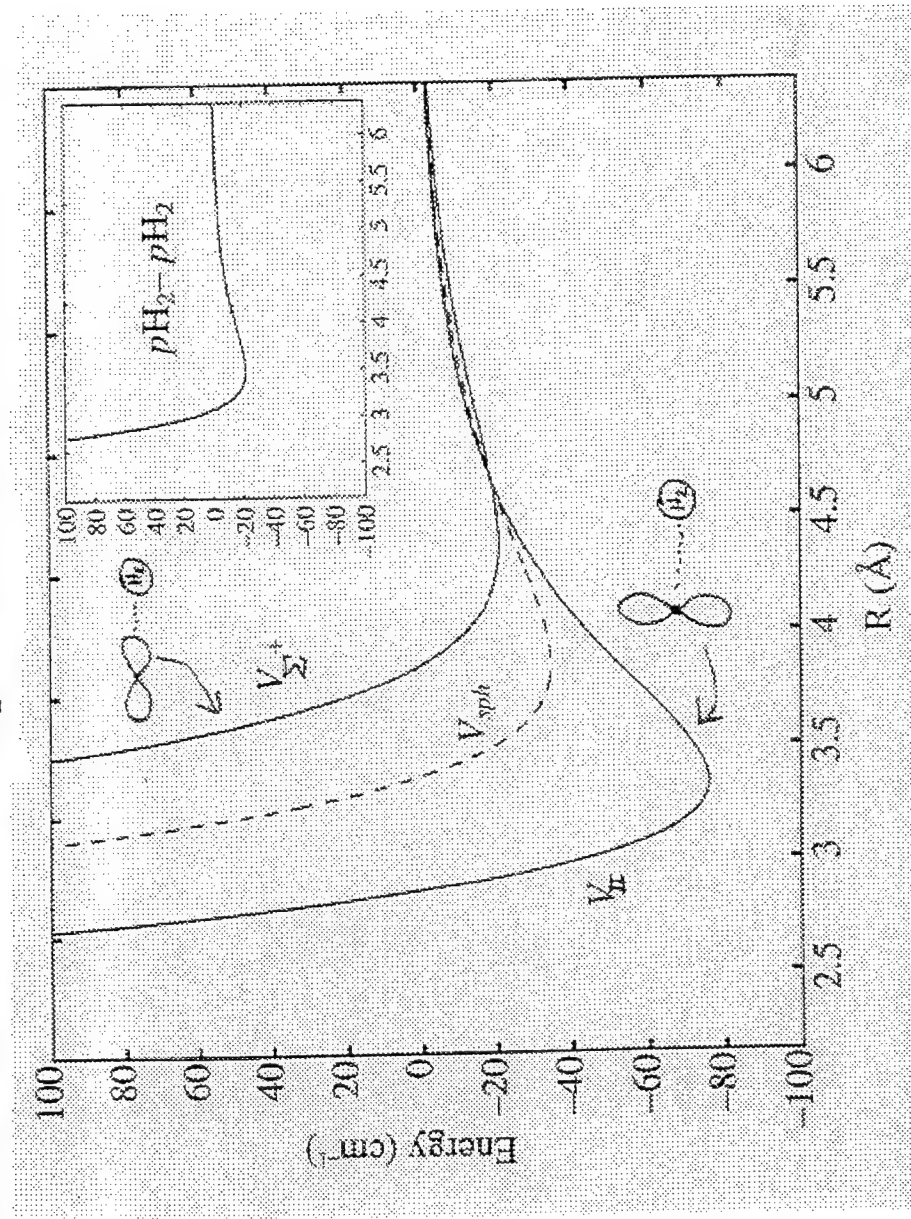
Computational requirements: ~50,000 CPU-hrs, MHPCC IBM SP

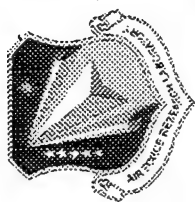


RESULTS - HEDM

Atom-doped solid hydrogen

How important is the orientational dependence of B-H₂ interactions in B/sH₂?

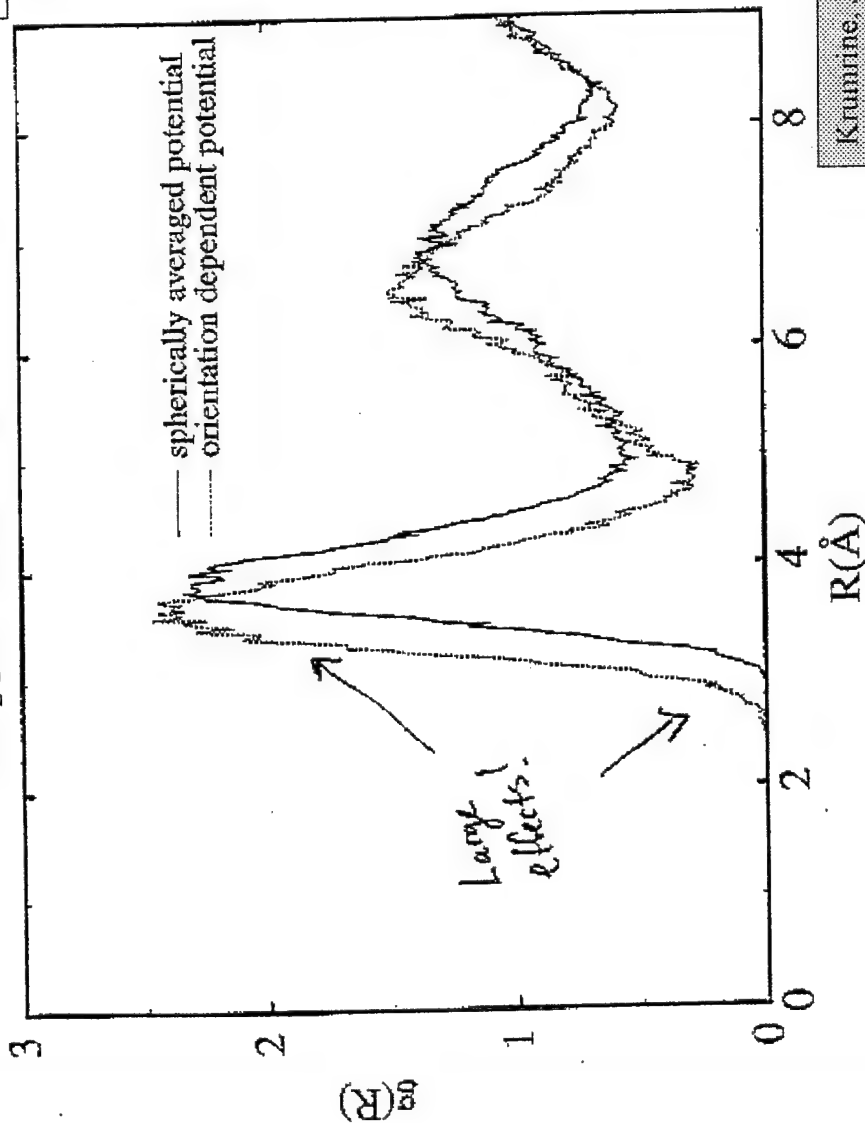




RESULTS - HEDM

Atom-doped solid hydrogen

B-H₂ pair correlation function



B-H₂ interaction energies (cm⁻¹)

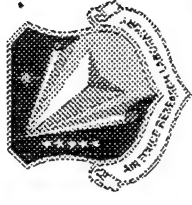
Site	$\langle V_{B,H_2} \rangle$
1-atom substitution	-546.72 (o.d.) -443.32 (sph.)
1-atom sub. + vacancy	-538.49 (o.d.) -439.88 (sph.)
surface	-375.32 (o.d.) -261.64 (sph.)

The bottom line: more accurate potentials predict greater stability of B atoms in solid H₂!

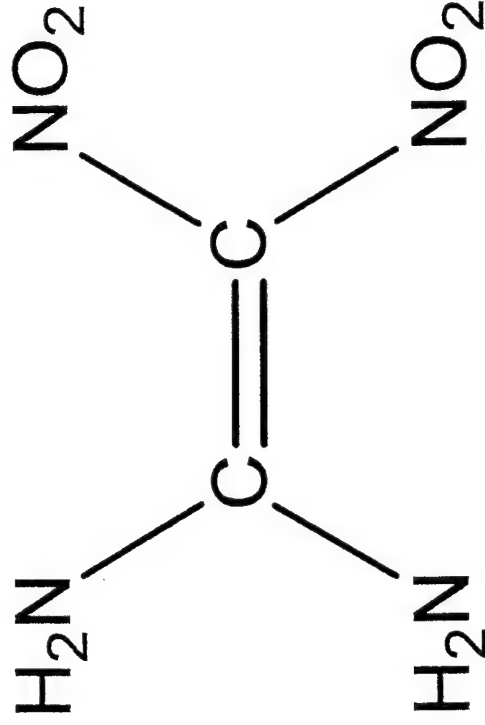
Krumrine, J.R., Jang, S., Alexander, M.H., and Voith,
G.A. J. Chem. Phys. 113 (2000) 9079



RESULTS - HEDM



FOX-7 (1,1-diamino-2,2-dinitroethylene)



$I_{sp} = 254$ sec (calculated)
 $\Delta H_f = -9.5$ kcal/mol (G2(MP2))

Advantages:

- Chemically balanced
wrt decomposition
products ($2CO + 2H_2O + 2N_2$)

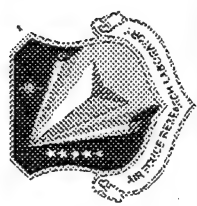
- Lower impact/shock
sensitivity than other

$C_nH_{2n}O_{2n}N_{2n}$ compounds
(e.g., RDX and HMX).

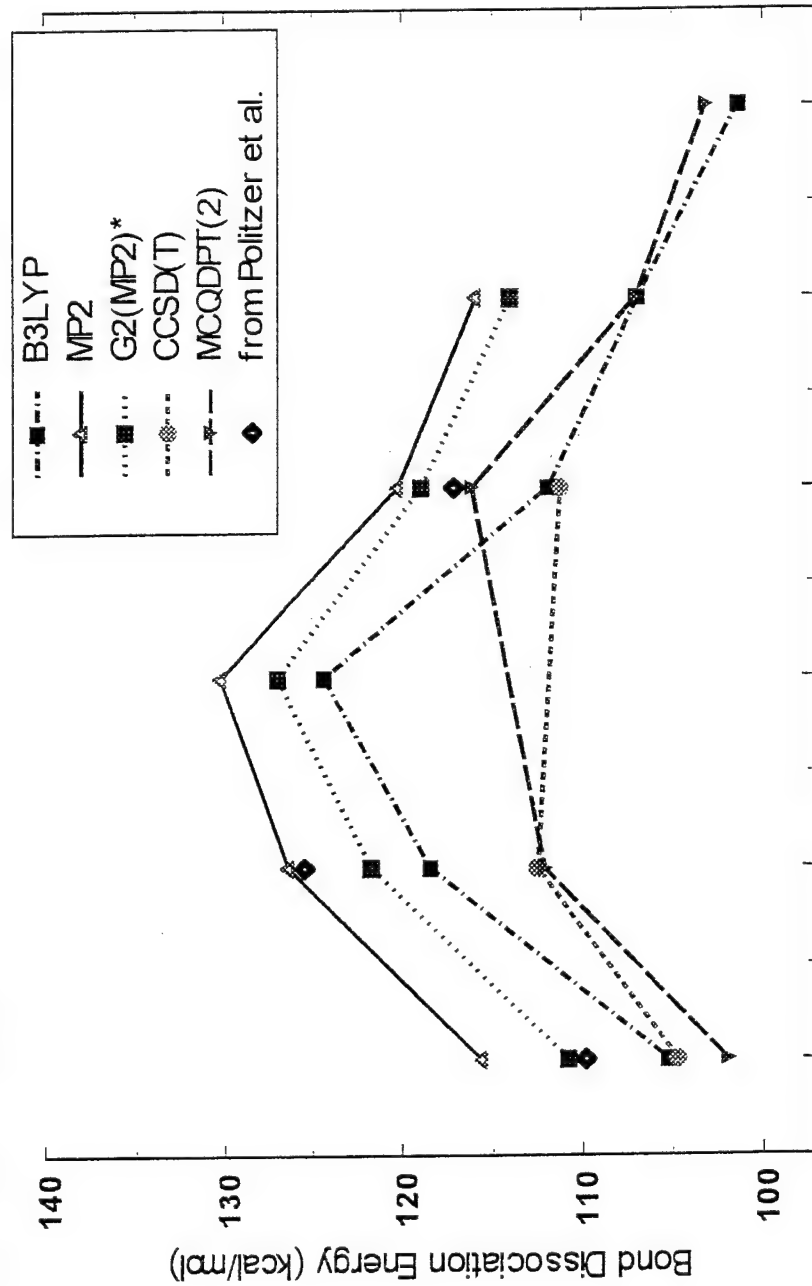
D.C. Sorescu, J.A. Boatz, D.L. Thompson, J. Phys. Chem. A, 105, 5010(2001).



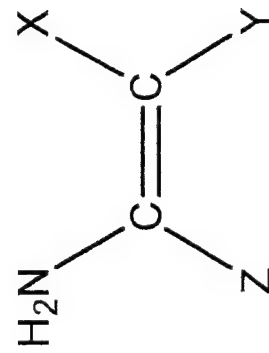
RESULTS - HEDM

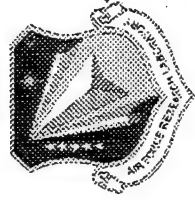


Development of a solid-state force field: calibration of theoretical methods for computing bond dissociation energies



HPC Requirements: 2 CPU-months, ARL SV-1, 1.3 GW memory, 190 GB disk



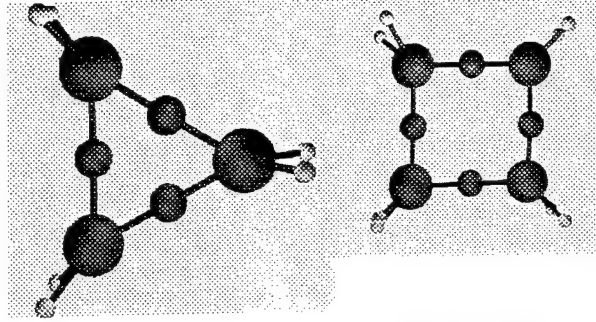


RESULTS - POSS

Mechanism of formation

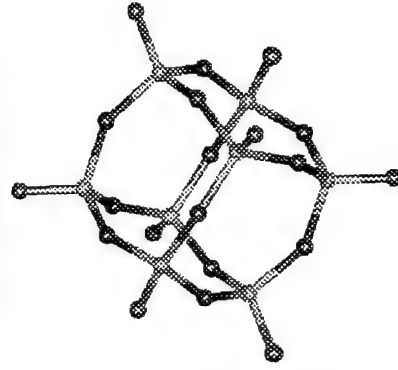
Key steps

1. Hydrolysis of RSiX_3 ($\text{R}=\text{H}, \text{CH}_3, \text{t-butyl}$, etc.; $\text{X}=\text{Cl}$)
 $\text{RSiCl}_3 + \text{H}_2\text{O} \rightarrow \text{RSiCl}_2\text{OH} + \text{HCl}$
 $\text{RSiCl}_2\text{OH} + \text{H}_2\text{O} \rightarrow \text{RSiCl(OH)}_2 + \text{HCl}$
 $\text{RSiCl(OH)}_2 + \text{H}_2\text{O} \rightarrow \text{RSi(OH)}_3$
2. Condensation of RSi(OH)_3 to disiloxane
 $2 \text{RSi(OH)}_3 \rightarrow \text{R(OH)}_2\text{SiOSi(OH)}_2\text{R} + \text{H}_2\text{O}$
3. Condensation of disiloxane to D_3, D_4
 $\text{RSi(OH)}_3 + \text{R(OH)}_2\text{SiOSi(OH)}_2\text{R} \rightarrow \text{D}_3 + 2\text{H}_2\text{O}$
 $[2+2]: 2\text{R(OH)}_2\text{SiOSi(OH)}_2\text{R} \rightarrow \text{D}_4 + 2\text{H}_2\text{O}$
 $[3+1]: \text{RSi(OH)}_3 + \text{R(OH)}_2\text{SiOSi(OH)}_2\text{R} \rightarrow \text{D}_4 + 2\text{H}_2\text{O}$
Ring Expansion: $\text{RSi(OH)}_3 + \text{D}_3 \rightarrow \text{D}_4 + \text{H}_2\text{O}$
4. Condensation of D_3, D_4 to POSS (in progress)
 $2\text{D}_4 \rightarrow \text{T}_8 + 4\text{H}_2\text{O}$
....

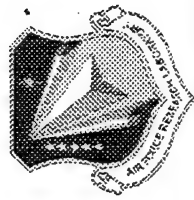


D3

D4



T8



RESULTS - POSS

Mechanism of formation: role of solvent (H_2O) & substituents (R)



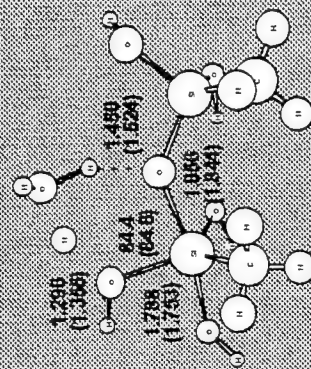
R Energy barrier (kcal/mol)

HF/6-31G* MP2/6-31G*

H	30.4 (16.7)	10.9 (-9.3)
Me	28.2 (14.7)	7.7 (-13.3)
t-Bu	34.3 (24.9)	9.8 (-9.3)
Ph	31.1 (18.2)	7.9 (-16.4)

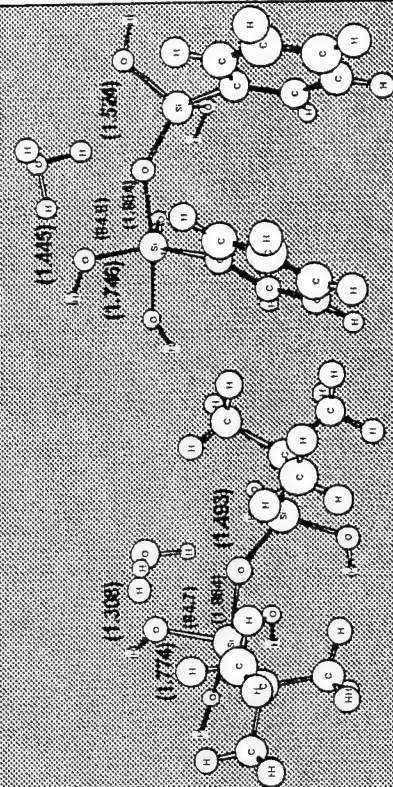
Values in parentheses are for water-catalyzed results.

TS for the condensation of RSi(OH)_3 catalyzed by a water



R=Me

MP2/6-31G*
(HF/6-31G*)



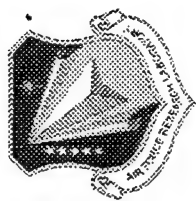
R=t-Bu

R=Ph

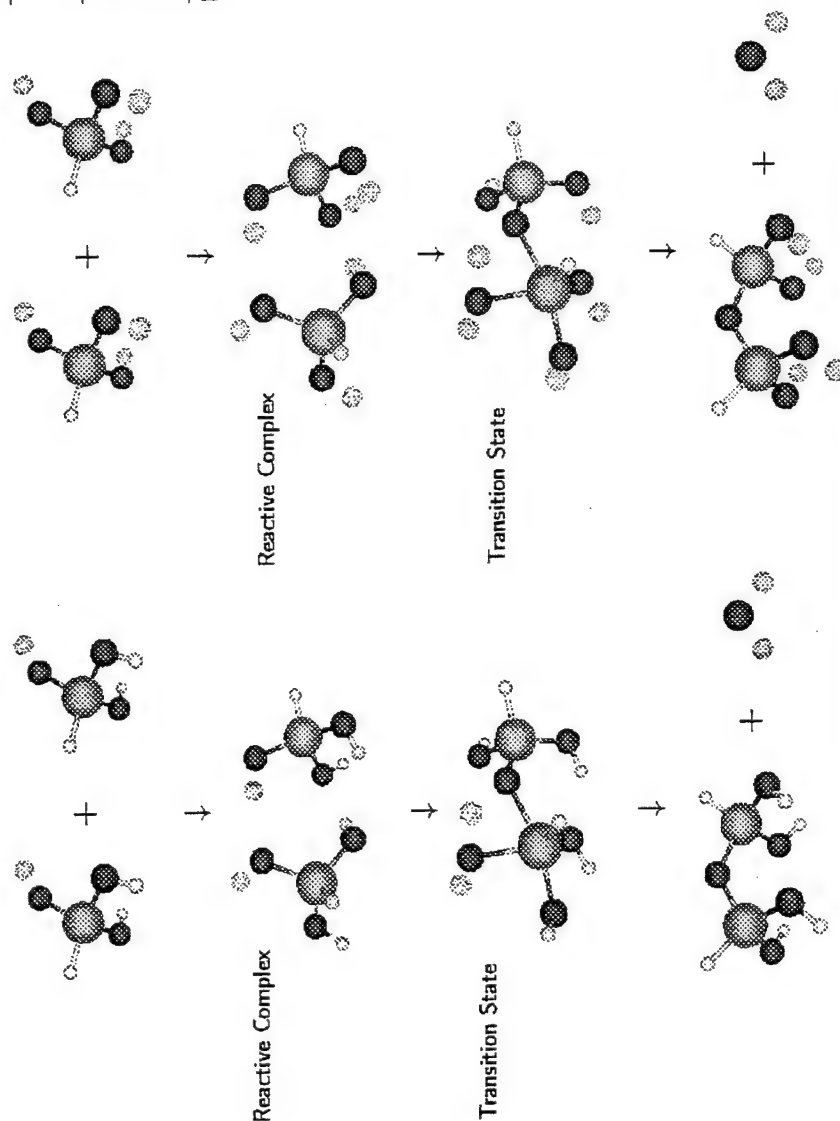
Kudo, T., Gordon, M.S. J. Am. Chem. Soc., 120, 11432 (1998)
Kudo, T., Gordon, M.S. J. Phys. Chem. A, 104, 4058 (2000)



RESULTS - POSS

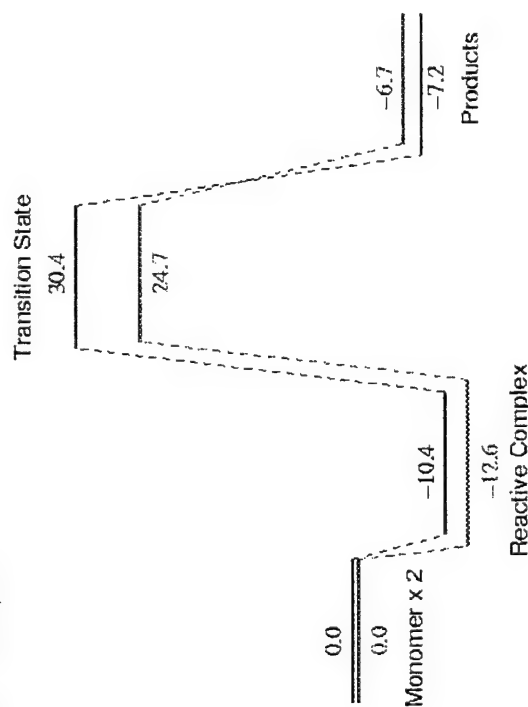


Nuclear quantum effects in proton transfer reactions



Level	Monomer $\times 2$	Reactive Complex	Transition State	Products	CPU Time
RHF/6-31G*	0.0	-10.4	30.4	-7.2	1.0
NEO-HF/2	0.0	-11.4	24.9	-5.9	1.1
NEO-HF/6	0.0	-12.6	24.7	-6.7	1.8

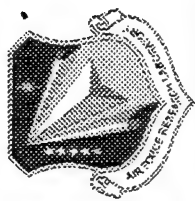
Units in kcal/mol



Hammes-Schiffer, S. J. Phys. Chem. A 102 (1998), 10443

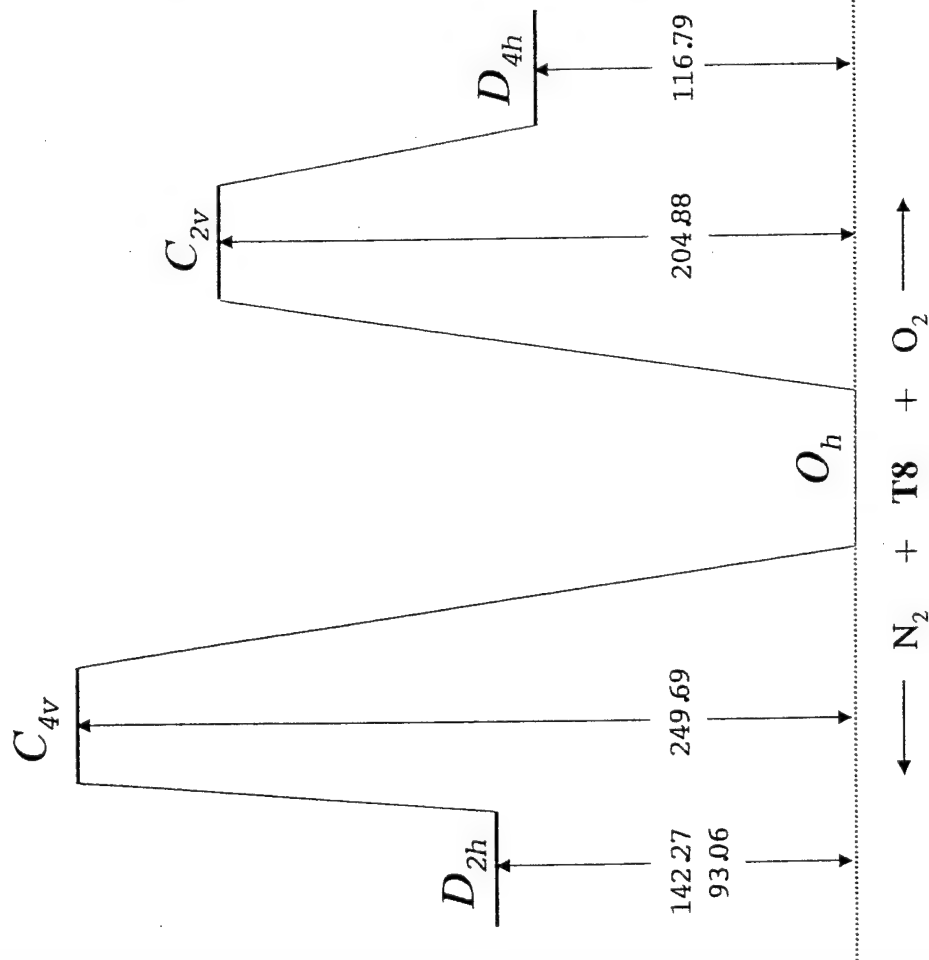
Webb, S.P., Agarwal, P.K., and Hammes-Schiffer, S. J. Phys. Chem. B, 104(2000), 888

Webb, S.P. and Hammes-Schiffer, S. J. Chem. Phys. 113 (2000), 5214

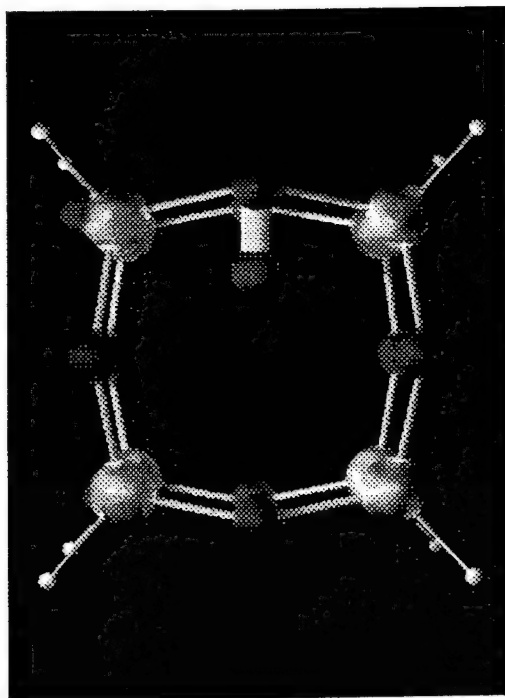


RESULTS - POSS

Molecular "sieves": preferential capture N_2 vs. O_2 ?



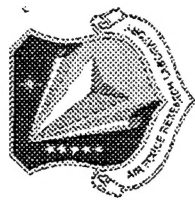
TS structure of $T_8 + O_2$



T_{10} and T_{12} calculations in progress
HPC requirements: ~50,000 node-hrs, AHPARC T3E, 256 GB



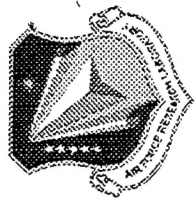
RESULTS - NLO



B3LYP S_0 - T_1 Excitation Energies (in eV)

System	6-31G(d)	Error	Exp
Porphyrin (1^3B_{2u})	1.42	0.16	1.58 ^a
Zinc Porphyrin (1^3B_{1u})	1.65	0.07	1.72 ^b
Tetraphenylporphyrin (1^3B_1)	1.31	0.14	1.45 ^c
Zinc Tetraphenylporphyrin(1^3B_1)	1.53	0.06	1.59 ^d
Zinc Phthalocyanine (1^3B_{2u})	1.05	0.08	1.13 ^e
Zinc Tetraabenzporphyrin (1^3B_{1u})	1.41	0.16	1.57 ^f
Phthalocyanine (1^3B_{1u})	1.18	0.06	1.24 ^g
Mean Error		0.10	

^aGouterman, Khalil, *J. Mol. Spectrosc.* 1974, 53, 88. (EPA (5:5:2) mixture of ethyl ether to isopentane to ethanol) and 50% ethyl iodide at 77 K) ^bGradyushko, Tsvirko, *Opt. Spectrosc.* 1971, 31, 291.(EPA at 77 K) ^cGouterman, Khalil, *J. Mol. Spectrosc.* 1974, 53, 88. (EPA at 77 K) ^dWalters et al., *J. Phys. Chem.* 1995, 99, 1166.(1:1 mixture of ether to ethanol at 77 K) ^eVincett et al.,K. E. *J. Chem. Phys.* 1971, 55, 4131. (1-chloronaphthalene at 77 K) ^fBajema, Gouterman, *J. Mol. Spectrosc.* 1971, 39, 421 (octane at 77 K) ^gMcVie et al., *J. Chem. Soc. Faraday Trans. II* 1978, 74, 1870 (1-chloronaphthalene at 77 K)



SUMMARY

High Energy Density Matter

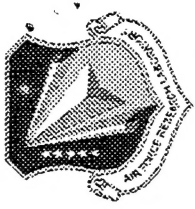
- High-nitrogen/polynitrogen compounds are more energetic than hydrazine.
- Inclusion of anisotropic interactions of B atoms in sH_2 predict greater stability than spherical interaction model.
- Extensive network of hydrogen bonds in FOX-7 may be responsible for lower shock, friction sensitivity relative to HMX, RDX.

Polyhedral Oligomeric Silsesquioxanes (POSS)

- Proton transfer reactions for hydrolysis and condensation are catalyzed by water.
- Alkyl substituents (R) in $RSiX_3$ have minor effects on hydrolysis and condensation reaction barriers.
- Nuclear quantum effects are important in proton transfer reactions -- lower barriers by $\sim x$ kcal/mol.
- T_g is too small to encapsulate N_2 or O_2 .



SUMMARY (cont.)

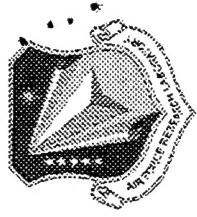


NLO materials

- Time-dependent density functional theory accurately predicts NLA in porphyrins.
- Computed triplet-triplet excitation energies within 0.1-0.4 eV of experiment
- Computed singlet-triplet excitation energies within 0.1-0.2 eV of experiment
- Computed ionization potentials accurate within 0.1 eV of experiment



ACKNOWLEDGEMENTS



POSS: Takako Kudo, Shawn Phillips, Simon Webb, Frank Feher, Joe Lichtenhan

HEDM: Millard Alexander, Jennifer Krumrine, Soomin Jang, Jeff Mills, Jeff Sheehy, Don Thompson, Dan Sorescu

NLO: Kiet Nguyen, Paul Day

MSRCs, DCs: ASC, ARL, ERDC, NAVO, MHPCC, AHPCRC

CHSSI funding (CCM-2, CCM-4)